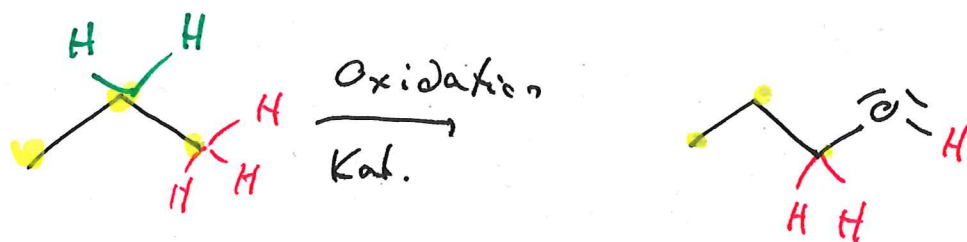
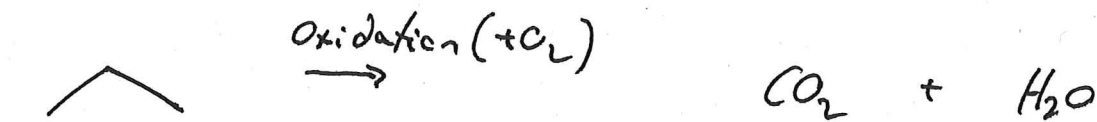


OC II

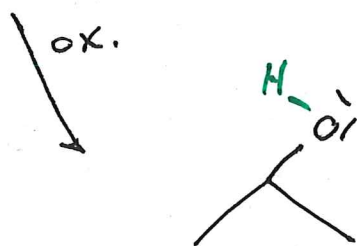
2 "Herstellen der diversen funktionellen Gruppen"

Redox → Oxidationszahlen

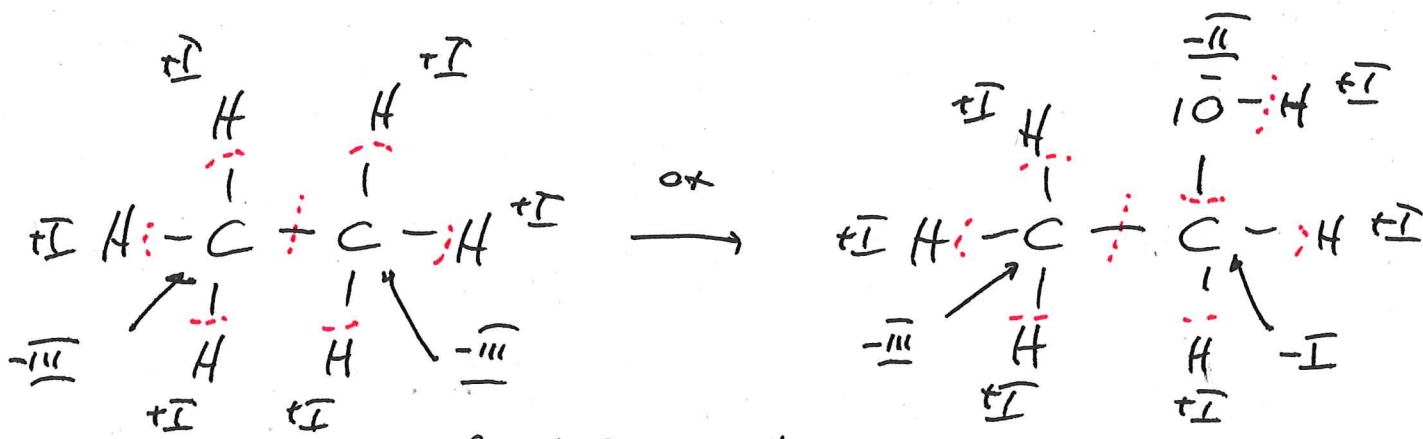
Oxidation :
 • Elektronenabgabe
 • Reaktion mit O_2



Alkan



Alkohol

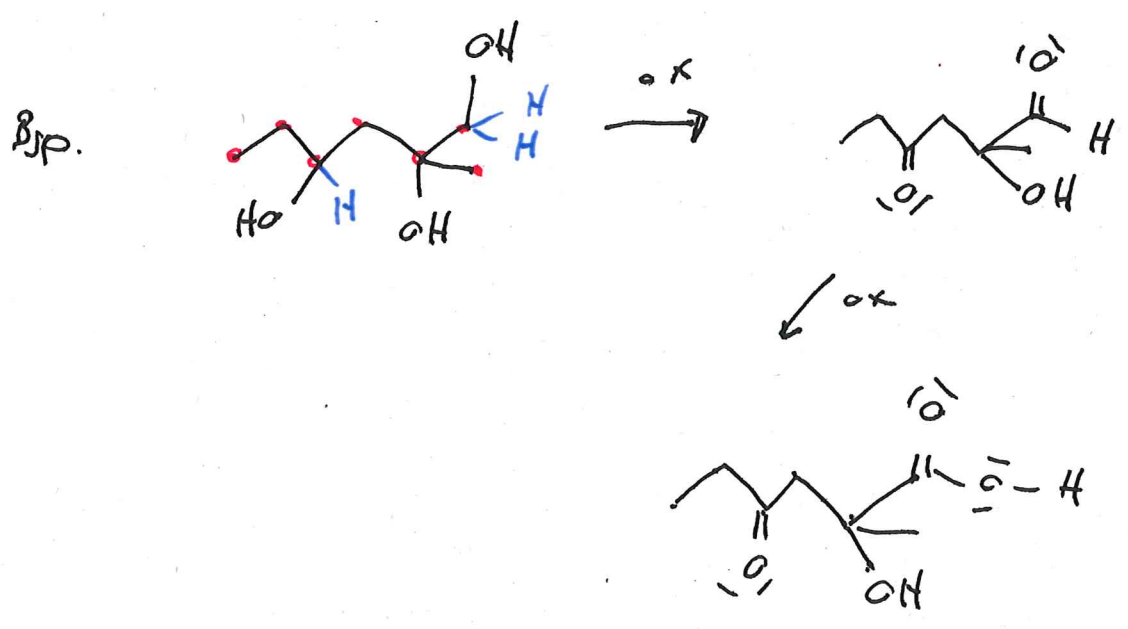
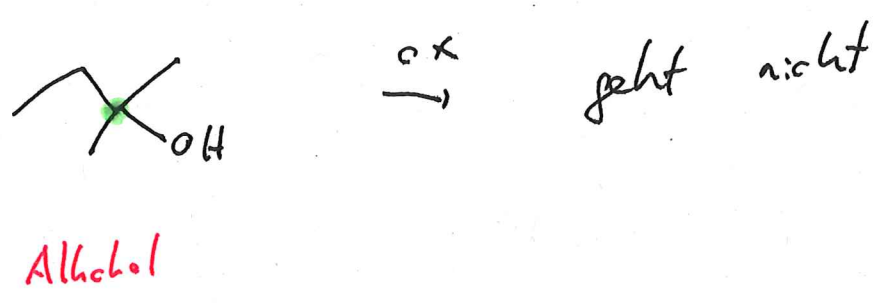
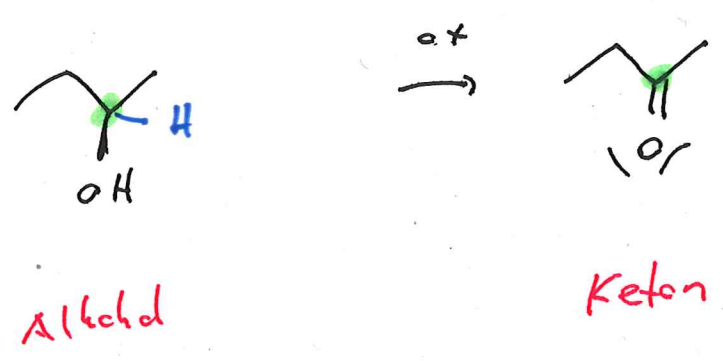
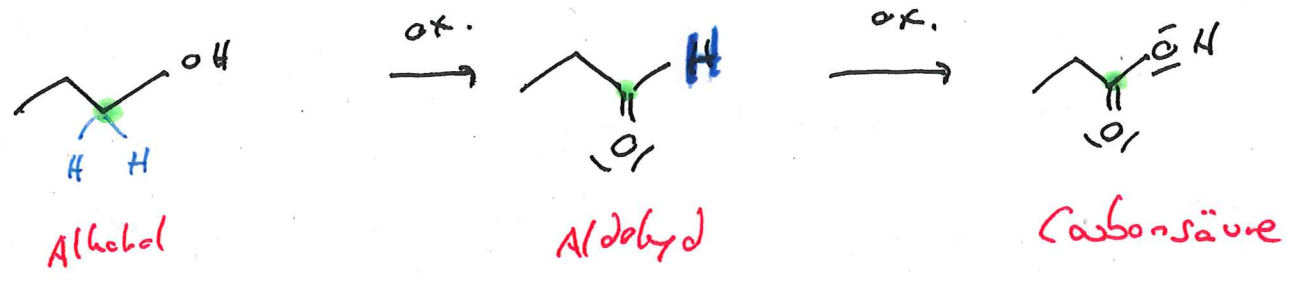


$EN(H) = 2.2$
 $EN(C) = 2.6$
 $EN(O) = 3.4$

$H \cdot$
 $\cdot C$
 $\cdot O$

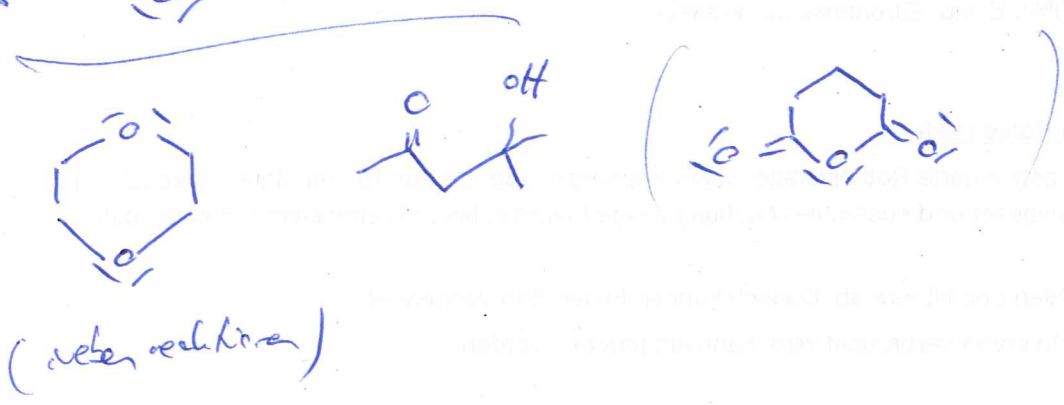
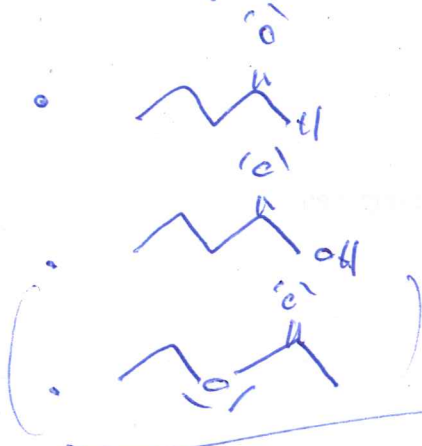
② OC II

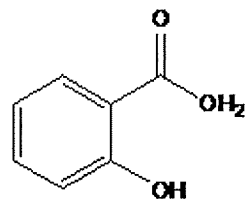
2.1. Weiteroxidation des Alkohols (Oxidationsstrasse)



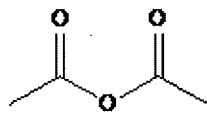
3. ab ...

stelle C_5 (Ausgangsaldehyd) sei jeweils ein
Alkanperoxid

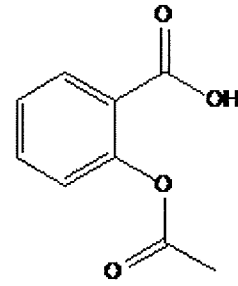
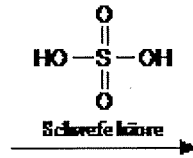




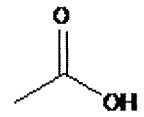
Salicylsäure

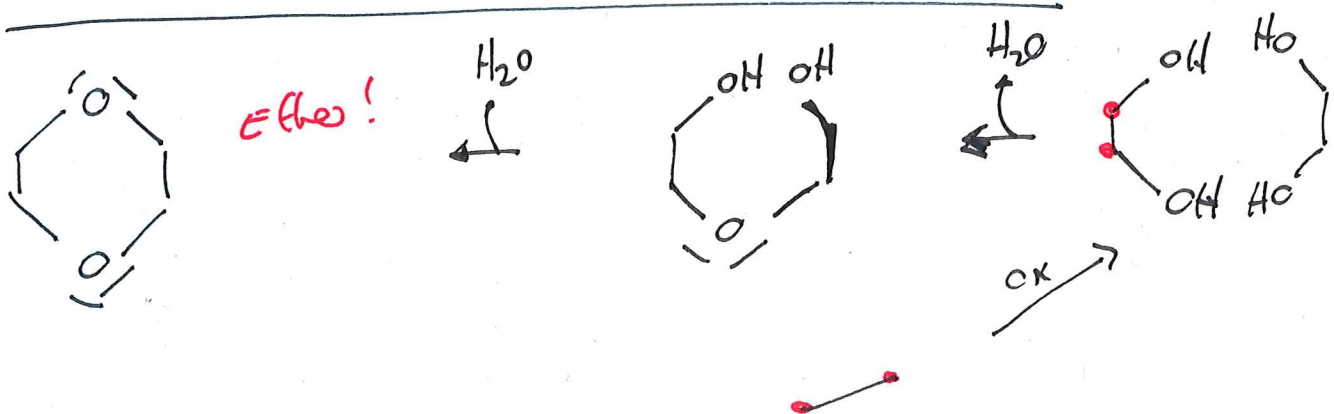
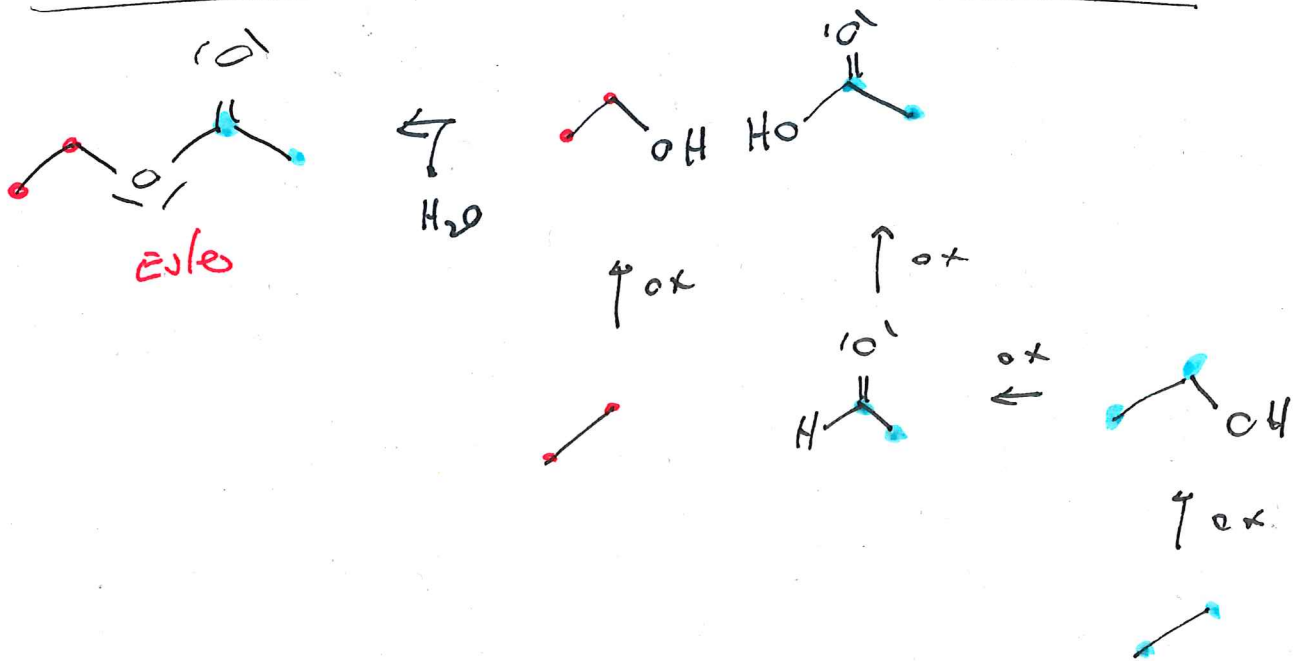
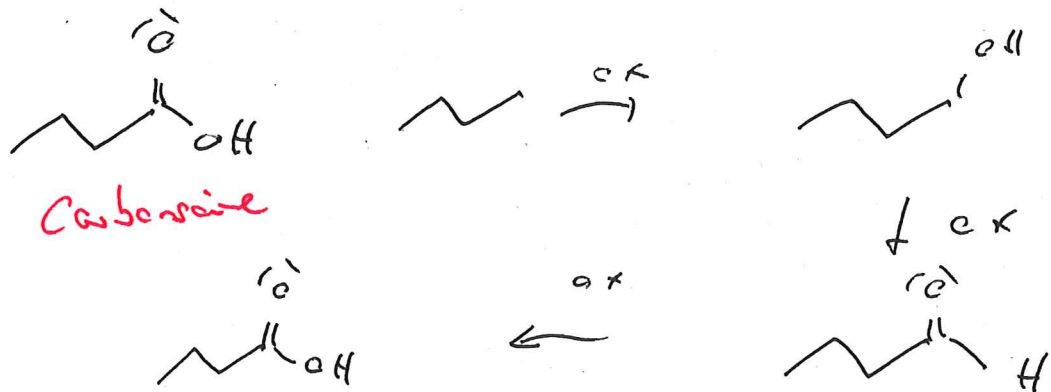


Essigsäureanhydrid

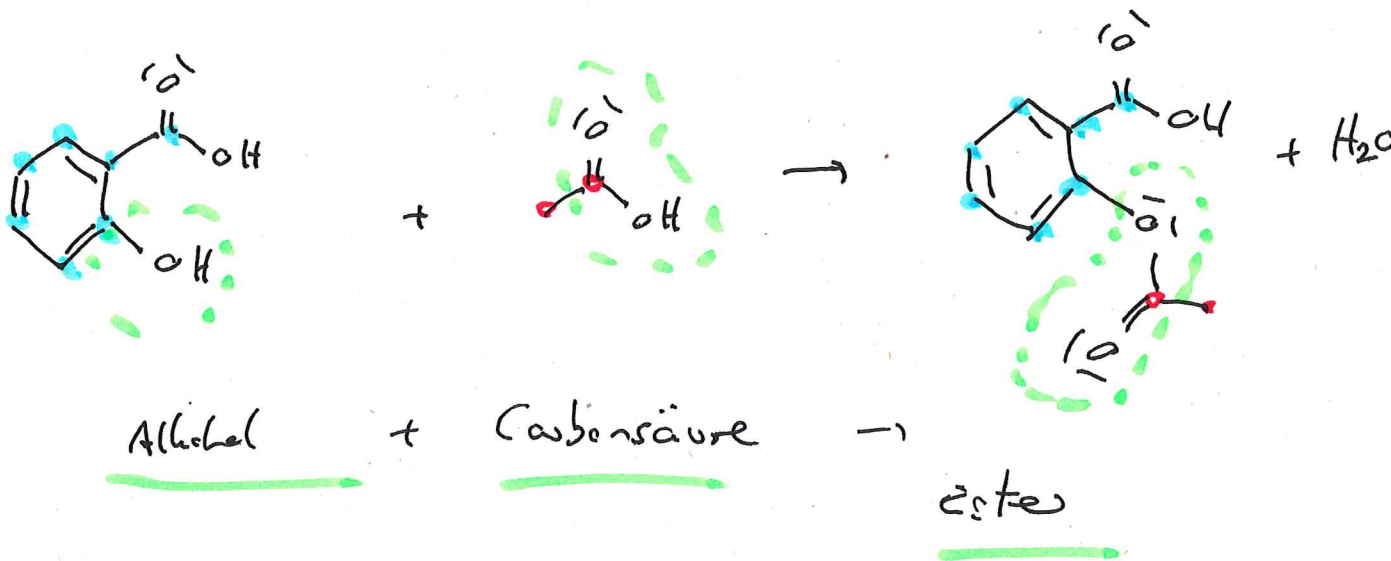


Acetylsalicylsäure





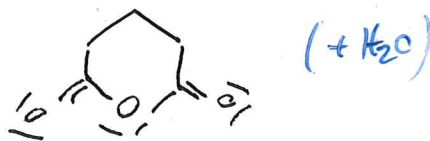
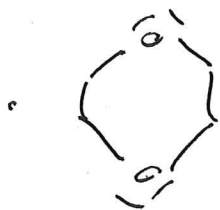
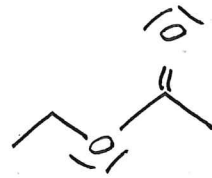
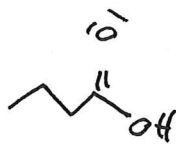
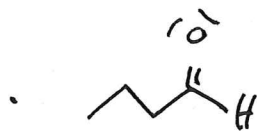
2.2. Herstellen eines Esters

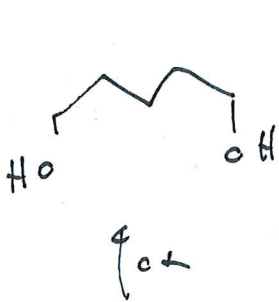
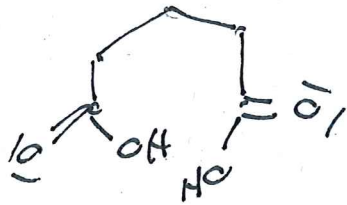
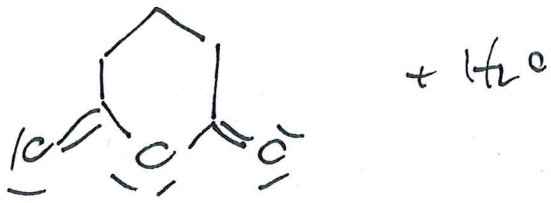


- vgl. Praktikum:
- Aspirin-Synthese
 - Wasserabweiser

2.3. Aufgaben

Stelle folgende Substanzen her, Ausgangsprodukt sei jeweils ein Alkanerüst.

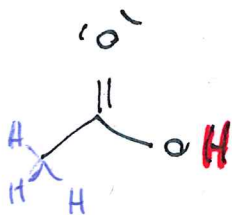




Carbonsäuren

$c = \text{mol/l}$

Essig-
säure



$$pH = -\log(c(\text{H}_3\text{O}^+))$$

Säure: protonenspendendes H^+



pH = 1
sauer



pH = 7 neutral

→ pH = 14
basisch